

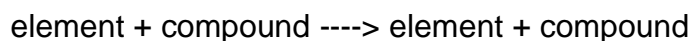
3. Single Displacement (also called Single Replacement) - In this type of reaction, a neutral element becomes an ion as it replaces another ion in a compound. The general form of this equation can be written as:



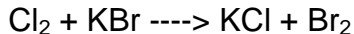
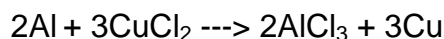
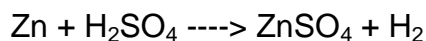
or



In either case we have:



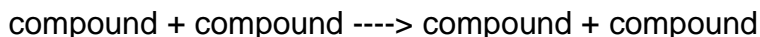
Some examples of single displacement reactions are shown below:



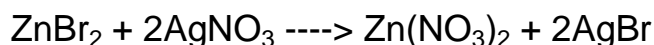
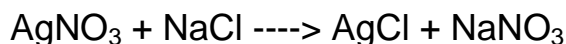
4. Double Displacement (also called Double Replacement) - Like dancing couples, the compounds in this type of reaction exchange partners. The basic form for this type of reaction is shown below:



or



Some examples of double displacement reactions are shown below;



5. Combustion - When organic compounds like propane are burned, they react with the oxygen in the air to form carbon dioxide and water. The reason why these combustion reactions will stop when all available oxygen is used up is because oxygen is one of the reactants. The basic form of the combustion reaction is shown below:

hydrocarbon + oxygen ----> carbon dioxide and water

Some examples of combustion reactions are;

